

1. Calculate the volume of a 0.200 M KOH solution that is needed to neutralize 25.00 mL of a 0.115 M HCl solution.

2. 34.62 mL of 0.1510 M NaOH was needed to neutralize 50.0 mL of an H_2SO_4 solution. What is the concentration of the original sulfuric acid (H_2SO_4) solution?

3. What volume of 0.075 M HCl is required to neutralize 100 ml of 0.01 M $\text{Ca}(\text{OH})_2$ solution?

4. A 25 ml solution of 0.5 M NaOH is titrated until neutralized into a 50 ml sample of HCl. What was the concentration of the HCl?

5. 30 mL of 0.10M NaOH neutralised 25.0mL of hydrochloric acid. What is the concentration of hydrochloric acid?

6. 50mL of 0.2mol/L NaOH neutralised 20mL of sulfuric acid. What is the concentration of the sulfuric acid?

7. 25.0mL of 0.05M Ba(OH)₂ neutralised 40.0mL of nitric acid. What is the concentration of the nitric acid?

8. If it takes 50 ml of 0.5M KOH solution to completely neutralize 125ml of sulfuric acid solution (H₂SO₄), what is the concentration of the H₂SO₄ solution?